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Chapter 10 - Chemical Calculations and Chemical Equations 141 The section ends with a summary of equation stoichiometry problems and shows how the skills developed in Section 10.1 can be mixed with the new skills developed in this section. Section 13.3 completes our process of describing equation stoichiometry problems by

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Chapter 10 Chemical Calculations And Equations

You will do separate calculations and add the results. General procedure: percent > grams of mixture > moles of mixture > moles of product > grams of product. -from percent to grams, use given percent in a ratio: given percent (as grams)/100.0 g. Sequential chemical reaction equation. Start with given information.

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370 Chapter 10 Chemical Calculations and Chemical Equations Example 10.1 shows how the coefficients in a balanced chemical equation provide a number of conversion factors that allow us to convert from moles of any reactant or product to moles of any other reactant or product. example 10.1 - Equation Stoichiometry

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Reactions of Chapter 10 Worksheet and Key A chemistry student working in the lab might be asked to calculate how much 1-bromo-2-methylpropane, C₄H₉Br, could be made from 6.034 g of ... 368 Chapter 10 Chemical Calculations and Chemical Equations. 10.1 Equation Stoichiometry 369 The ratio of moles of P₄O₁₀ to moles of P ...

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370 Chapter 10 Chemical Calculations and Chemical Equations Example 10.1 shows how the coefficients in a balanced chemical equation provide a number of conversion factors that allow us to convert from moles of any reactant or product to moles of any other reactant or product. example 10.1 - Equation Stoichiometry

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up. down. 12 g of carbon makes 44 g of carbon dioxide. 12 ÷ 44 g of carbon will make 1 g of carbon dioxide. You will need 11 × (12 ÷ 44) = 3 g of carbon to make 11 g of carbon dioxide.

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Chapter 10 159 Exercises Key Exercise 10.1 - Equation Stoichiometry: Tetrachloroethene, C₂Cl₄, often called perchloroethylene (perc), is a colorless liquid used in dry cleaning. It can be formed in several steps from the reaction of dichloroethane, chlorine gas, and oxygen gas. The equation for the net reaction is

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