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two characteristics of
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and Pd form inert
electrodes, i.e., they are
not attacked by the ions
of the electrolyte or the
products of electrolysis.
Hence, they are used as
electrodes for carrying out

electrolysis.NCERT
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Chemistry Chapter 5
Surface ...3.7 Define
conductivity and molar
conductivity for the
solution of an electrolyte.
Discuss their variation
with concentration. Sol.
Conductivity of a solution
is defined as the
conductance of a solution
of 1 cm length and having
1 sq. cm as the area of
cross-section. Molar
conductivity of a solution
at a dilution (V) is the
conductance of all the
ions produced from one
mole of the electrolyte

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3.7 Define conductivity and molar conductivity for the solution of an

electrolyte. Discuss their variation with concentration. Sol. Conductivity of a solution is defined as the conductance of a solution of 1 cm length and having 1 sq. cm as the area of cross-section. Molar conductivity of a solution at a dilution (V) is the conductance of all the ions produced from one mole of the electrolyte ...
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Surface ...

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QUESTIONS SOLVED. 5.1.

Write any two characteristics of Chemisorption. Ans: Pt and Pd form inert electrodes, i.e., they are not attacked by the ions of the electrolyte or the products of electrolysis. Hence, they are used as electrodes for carrying out electrolysis.

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in the 12th standard,
chemistry is the most
used subject in our daily
lives. We find the use and
application of chemistry in
our everyday lives. From
medical studies to most
intricate gene studies we
will find that chemistry is
used everywhere.